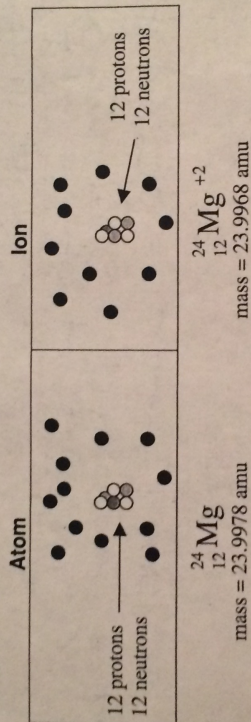
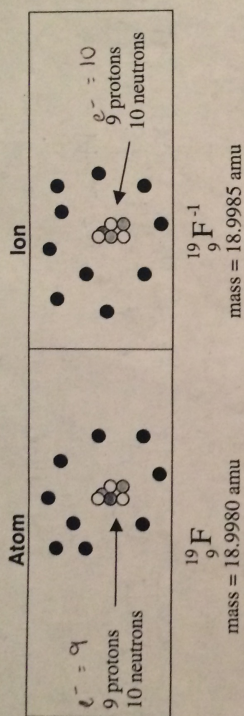


Information: Atoms, Ions, Masses of Subatomic Particles

The atomic mass unit (amu) is a special unit for measuring the mass of very small particles such as atoms. The relationship between amu and grams is the following: $1.00 \text{ amu} = 1.66 \times 10^{-24} \text{ g}$. Note the following diagrams comparing atoms and ions.

**Critical Thinking Questions**

10. What is structurally different between an atom and an ion? Note: This is the ONLY structural difference between an atom and an ion.

Electron's different number

electrons = # protons
ion = different
more or less electrons than protons

11. In atomic mass units (amu), what is the mass of an electron?

less than one amu
0.0005 amu

0 for our purposes.

12. Is most of the mass of an atom located in the nucleus or outside the nucleus? How do you know?

In the nucleus because the mass numbers roughly equal the number of protons and neutrons.

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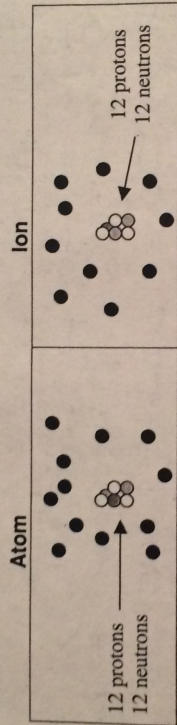
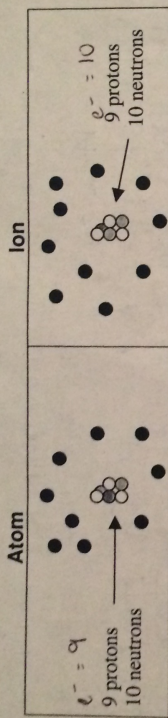


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Critical Thinking Questions

10. What is structurally different between an atom and an ion?
 difference between an atom and an ion.
 Electron's different.
 number

→ # electrons = # protons
 Note: This is the ONLY structural difference between an atom and an ion.
 MORE OR LESS electrons than protons
 → ion = different

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